

- (10) M. K. Jain, I. Kirson, and D. Lavie, *Isr. J. Chem.*, **4**, 32 (1966).
- (11) C. J. Breen, E. Ritchie, W. T. L. Sidwell, and W. C. Taylor, *Aust. J. Chem.*, **19**, 455(1966).
- (12) H. K. Adam, T. A. Bryce, I. M. Campbell, N. J. McCorkindale, A. Gaudemer, R. Gmelin, and J. Polonsky, *Tetrahedron Lett.*, **1967**, 1461.
- (13) A. Chatterjee and A. B. Kundu, *ibid.*, **1967**, 1471.
- (14) J. D. Connolly, K. L. Handa, R. McCrindle, and K. H. Overton, *ibid.*, **1967**, 3449.
- (15) D. Lavie, M. K. Jain, and I. Kirson, *J. Chem. Soc. (C)*, **1967**, 1347.
- (16) T. Murata, M. Shinohara, T. Hirata, K. Kamiya, M. Nishikawa, and M. Miyamoto, *Tetrahedron Lett.*, **1968**, 103.
- (17) T. Murata, M. Shinohara, T. Hirata, and M. Miyamoto, *ibid.*, **1968**, 849.
- (18) D. Lavie and E. C. Levy, *ibid.*, **1968**, 2097.
- (19) L. L. Smith, W. S. Matthews, J. C. Price, R. C. Bachmann, and B. Reynolds, *J. Chromatogr.*, **27**, 187(1967).
- (20) J. E. van Lier and L. L. Smith, to be published.
- (21) J. E. van Lier and L. L. Smith, *Biochemistry*, **6**, 3269(1967).
- (22) J. E. van Lier and L. L. Smith, *Anal. Biochem.*, **24**, 419 (1968).
- (23) J. E. van Lier and L. L. Smith, *Tex. Rep. Biol. Med.*, **27**, 167(1969).
- (24) P. Kurath, F. M. Ganis, and M. Radakovich, *Helv. Chim. Acta*, **40**, 933(1957).
- (25) N. Entwistle and A. D. Pratt, *Tetrahedron*, **25**, 1449(1969).
- (26) D. I. Duveen and J. Kenyon, *J. Chem. Soc.*, **1936**, 1451.
- (27) S. R. Landor, B. J. Miller, and A. R. Tatchell, *J. Chem. Soc. (C)*, **1966**, 1822, 2280.
- (28) P. A. Plattner, *Helv. Chim. Acta*, **45**, 1693(1951); *Chem. Ind.*, **1951**, SNI.
- (29) W. Klyne and W. M. Stokes, *J. Chem. Soc.*, **1954**, 1979.
- (30) K. Tsuda and R. Hayatsu, *J. Amer. Chem. Soc.*, **81**, 5987 (1959).
- (31) H. Mori, K. Shibata, K. Tsuneda, M. Sawai, and K. Tsuda, *Chem. Pharm. Bull. Tokyo*, **16**, 1407(1968).
- (32) H. Mori, K. Shibata, K. Tsuneda, and M. Sawai, *ibid.*, **16**, 2416(1968).
- (33) E. P. Burrows, G. M. Hornby, and E. Caspi, *J. Org. Chem.*, **34**, 103(1969).
- (34) A. Ercoli, S. Di Frisco, and P. de Ruggieri, *Gazz. Chim. Ital.*, **83**, 78(1953).
- (35) A. Ercoli and P. de Ruggieri, *ibid.*, **83**, 720(1953).
- (36) A. Ercoli and P. de Ruggieri, *J. Amer. Chem. Soc.*, **75**, 3284(1953).
- (37) K. Mislow, "Introduction to Stereochemistry," W. A. Benjamin, New York, N. Y., 1966, p. 95.
- (38) T. Masui and E. Staple, *Steroids*, **9**, 443(1967).
- (39) W. v. E. Doering and R. W. Young, *J. Amer. Chem. Soc.*, **74**, 2997(1952).

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New Compounds: Mannich Bases from 1,2-Diphenylindolizine—*N*-Substituted Cyclohexylaminomethyl Derivatives

WILLIAM B. HARRELL*, SHAO-WEN KUANG, and CROSSLEY O'DELL

Abstract □ Seven new Mannich bases, involving *N*-substituted cyclohexylamines and 1,2-diphenylindolizine, have been synthesized as potential biologically active compounds.

Keyphrases □ Mannich bases—synthesis □ *N*-Substituted cyclohexylamines—synthesis from 1,2-diphenylindolizine □ 1,2-Diphenylindolizine—Mannich base synthesis

In previously reported work, it has been shown that certain Mannich bases derived from indolizines exhibited CNS-depressant activity (1–3). As part of a continuing exploration of indolizines with potential biological activity, a series of Mannich bases involving *N*-sub-

stituted cyclohexylamines was synthesized from 1,2-diphenylindolizine (4) (Table I).

EXPERIMENTAL¹

The appropriate secondary amine (0.045 mole) was combined with 30 ml. of 1,4-dioxane and 2.25 ml. of 40% aqueous formaldehyde (0.030 mole). The mixture was placed in the refrigerator and allowed to stand for 48 hr. To the mixture was then added 4.1 g. of 1,2-diphenylindolizine (0.015 mole); the resulting clear solution was stirred at room temperature for 72 hr., during which

¹ Melting points were taken on a Thomas-Hoover melting-point apparatus and are uncorrected. Elemental analyses were obtained from Strauss Microanalytical Laboratories, Oxford, England.

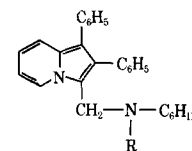


Table I—Mannich Bases

Compound	R	M.p.	Yield, %	Formula	Anal.	
					Calcd.	Found
I		182–183°	85	C ₃₃ H ₃₈ N ₂	C, 85.67 H, 8.28 N, 6.05	C, 85.72 H, 8.32 N, 6.18
II	CH ₃ CH ₂ —	123–124°	90	C ₂₉ H ₃₃ N ₂	C, 85.25 H, 7.89 N, 6.86	C, 85.32 H, 7.92 N, 6.69
III	CH ₃ —	112–113°	88	C ₂₈ H ₃₀ N ₂	C, 85.24 H, 7.66 N, 7.10	C, 85.36 H, 7.66 N, 7.09
IV	(CH ₃) ₂ CH—	127–128°	84	C ₃₀ H ₃₄ N ₂	C, 85.26 H, 8.11 N, 6.63	C, 85.42 H, 8.09 N, 6.42
V	NCCH ₂ CH ₂ —	152–153°	92	C ₃₀ H ₃₁ N ₃	C, 83.10 H, 7.21 N, 9.69	C, 83.13 H, 7.19 N, 9.48
VI	HOCH ₂ CH ₂ —	138–139°	74	C ₂₉ H ₃₂ N ₂ O	C, 82.04 H, 7.60 N, 6.60	C, 82.11 H, 7.65 N, 6.78
VII	-CHOHCH ₂ —	137–138°	84	C ₃₅ H ₃₆ N ₂ O	C, 83.96 H, 7.25 N, 5.59	C, 83.92 H, 7.31 N, 5.42

the crystalline product appeared. The product was collected and recrystallized from acetone. Each compound gave a negative color reaction with *p*-dimethylaminobenzaldehyde reagent, indicating that substitution had occurred at the C-3 position (5).

REFERENCES

- (1) W. B. Harrell and R. F. Doerge, *J. Pharm. Sci.*, **56**, 225(1967).
- (2) W. B. Harrell and R. F. Doerge, *ibid.*, **56**, 1200(1967).
- (3) W. B. Harrell and R. F. Doerge, *ibid.*, **57**, 1989(1968).

- (4) P. A. Barrett, *J. Chem. Soc.*, **1958**, 325.
- (5) D. O. Holland and J. H. C. Naylor, *ibid.*, **1955**, 1657.

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COMMUNICATIONS

Time Integral of Drug Concentration in the Central (Plasma) Compartment

Keyphrases Drug concentration time integral—central compartment Absorption comparison method—different formulations

Sir:

A well-known result in pharmacokinetics is that the time integral of the concentration of a drug in the central (plasma) compartment is equal to the total amount of drug absorbed divided by the product of the volume of distribution for the compartment and the elimination rate constant. The result is of great practical importance in that amounts of drug absorbed from different formulations of a drug can be readily compared by administering the different formulations to the same subject. Standard statistical designs, such as

balanced incomplete block designs, can thus be employed. The result has been proved for one-compartment and two-compartment systems under suitable conditions. The usual procedure has been to obtain an expression for the concentration in the central compartment (by solving the appropriate differential equations) and to integrate this expression over time to obtain the stated result. A recent example of this procedure is given in Eqs. 11a through 14a of Gibaldi *et al.* (1). The purpose of the present note is to show that the result is a direct consequence of two basic assumptions and thus holds under quite general conditions. In fact, the present proof is implicit in the derivation given for nonintravenous routes of administration in Eq. 22a of the reference.

The two basic assumptions are:

1. Elimination of the drug takes place only from the central compartment, that is, the compartment over